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Practice w/ Balancing Equations

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1. First, let's balance the calcium atoms by placing a three in front of CaCl_2 : $3\text{CaCl}_2 + \text{Na}_3\text{N}_2 \rightarrow 3\text{CaCl}_2 + \text{Na}_3\text{N}_2$. 2. Next, let's balance the nitrogen atoms, by placing a 2 in front of Na_3N : $3\text{CaCl}_2 + 2\text{Na}_3\text{N} \rightarrow 3\text{CaCl}_2 + 2\text{Na}_3\text{N}$. 3. Now we will balance the sodium atoms by placing a 6 in front of NaCl .

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Chemquest 29 Balancing Equations Answer Key chemquest 29 balancing equations answer. 1. First, let's balance the

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calcium atoms by placing a three in front of CaCl_2 : $\text{Ca}_3\text{N}_2 + \text{NaCl} \rightarrow 3\text{CaCl}_2 + \text{Na}_3\text{N}$. Next, let's balance the nitrogen atoms, by placing a 2 in front of Na_3N : $\text{Ca}_3\text{N}_2 +$

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3 a) Alzo: 3 u): Nuro lte c) 4 NazSo+ ,3C al 5 Mgc. {o:}z. Information: How To Balance Equations. Consider the reaction of sodium and chlorine to produce sodium chloride from question one: $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$. Remember that when chlorine is by itself it is always written as Cl_2 .

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the same exponents. Your answer will them have the same exponent as the numbers you add or. subtract. Here are some examples. Make sure you understand each step: $4.2 \times 10^6 + 3.1 \times 10^5$ make exponents the same, either a 5 or 6 $42 \times 10^5 + 3.1 \times 10^5 = 45.1 \times 10^5 = 4.51 \times 10^6$

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1. Balanced equations. (Coefficients equal to one (1) do not need to be shown in your answers). (a) $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$. (b) $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. (c) $2\text{FeBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 1\text{Fe}_2(\text{SO}_4)_3 + 6\text{HBr}$ (d) $1\text{C}_4\text{H}_6\text{O}_3 + 1\text{H}_2\text{O} \rightarrow 2\text{C}_2\text{H}_4\text{O}_2$. (e) $1\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ (f) $1\text{C}_4\text{H}_{10}\text{O} + 6\text{O}_2 \rightarrow 4\text{CO}_2 + 5\text{H}_2\text{O}$ (g) $1\text{C}_7\text{H}_{16} + 11\text{O}_2 \rightarrow 7\text{CO}_2 + 8\text{H}_2\text{O}$ (h) $4\text{H}_2\text{SiCl}_2 + 4\text{H}_2\text{O} \rightarrow 1\text{H}_8\text{Si}_4\text{O}_4 + 8\text{HCl}$ (i) $10\text{HSiCl}_3 + 15\text{H}_2\text{O} \rightarrow 1\text{H}_{10}\text{Si}_{10}\text{O}_{15} + 30\text{HCl}$ (j) $1\text{C}_7\text{H}_9 + 3\text{HNO}_3 \rightarrow 1\text{C}_7\text{H}_6(\text{NO}_2)_3 + 3\text{H}_2\text{O}$ (k) $1\text{C}_5\text{H}_8\text{O}_2 + 2\text{NaH} + 2\text{HCl} \dots$

Balancing Equations: Practice Problems

$\text{Cu}(\text{NO}_3)_2 + \text{Zn} \rightarrow \text{Zn}(\text{NO}_3)_2 + \text{Cu}$. There is a sentence that you can write to describe what happens above: Copper(II) nitrate reacts with zinc metal and yields zinc nitrate and copper metal. Take a minute to notice the relationship between the sentence and the symbols in the reaction. Now, try a few yourself.

Intro, to chemical Reactions - LTHS Answers

Hint: add your answer to question 1 and four times your answer to question 2. $4 + 4(7) = 4 + 28 = 32$. 4. Verify that there are 32 electrons pictured in the drawing of CCl_4 . Count each dot in the picture above: there are 24. There are also 8 more electrons represented by the 4 lines (each line represents two electrons).

ChemQuest 23 : Drawing Covalent Compounds

balancing equations answers key. The first step to balance the equation is to write down the chemical formula of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical equation.

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